

Chapter 1 and 2 Review

Give the number of sig figs for each value:

45.07

0.00356

7000

700.

100.0

3.00×10^5

9659

6.0×10^{-9}

Calculate the following problems:

$35.0\text{mL} + 36.65\text{mL} + 0.78\text{mL}$

$6.0\text{g} \times 10.3 \times 4.57$

A rock has a mass of 245g and a volume of 60cm^3 . What is its density?

An object with a mass of 25.3g is placed in water. The volume of the water started at 15.0 mL. When the object was added, the ending volume was measured at 18.0 mL. What is the density of the object?

The density of acetone is 0.784g/mL. If acetone has a mass of 142g, what is its volume of the liquid in liters?

Copper-63 has an atomic mass of 62.93 amu and a percent abundance of 30.9%. Copper-65 has an atomic mass of 64.93 amu with a percent abundance of 69.1%. What is the average atomic mass?

Determine the number of protons, neutrons, and electrons in the following elements:



Name the following compounds:



Give the chemical formula for the following compounds:

sodium chloride

potassium phosphate

nickel (II) sulfate

lithium sulfide

diphosphorus pentasulfide

chlorine trifluoride

copper (II) carbonate

chloric acid

nitrous acid